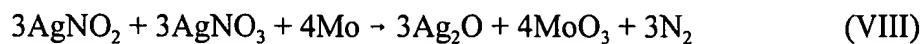


In the Specification:

Amend as follows at page 18:

Example 3.



C/ A comelt of equimolar amounts of AgNO_2 and AgNO_3 was mixed with a stoichiometric amount of Mo metal in accordance with equation VIII, i.e., 34.1% by weight AgNO_2 , 37.6% by weight AgNO_3 , and 28.3% by weight Mo. An autoignition temperature of $131 \pm 2^\circ\text{C}$ was determined for the composition using DSC.

✓ Example 4:



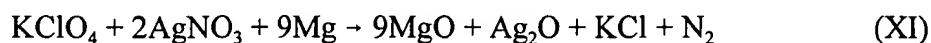
~~—————Lithium perchlorate, LiClO_4 , was mixed with a stoichiometric amount of Mo in accordance with equation IX, i.e., 45.4% by weight LiClO_4 and 54.6% by weight Mo. An autoignition temperature of $147 \pm 2^\circ\text{C}$ was determined for the composition using DSC.~~

C² Example 5.



AgNO_3 was mixed with a stoichiometric amount of magnesium, Mg, metal in accordance with equation X, i.e., 73.7% by weight AgNO_3 and 26.3% by weight Mg. An autoignition temperature of $157 \pm 2^\circ\text{C}$ was determined for the composition using DSC.

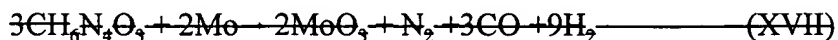
C³ Example 6.



Amend as follows at page 21:

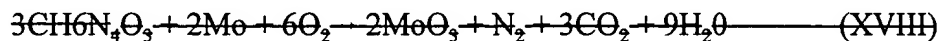
CH
metal in accordance with equation XVI, i.e., 20.5% by weight NaNO_3 , 41.0% by weight AgNO_3 and 38.5% by weight Mo. The composition autoignited at $217 \pm 2^\circ\text{C}$ by DSC analysis.

✓
Example 12:



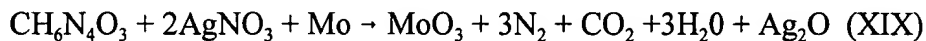
~~_____ Guanidine nitrate, $\text{CH}_6\text{N}_4\text{O}_3$, was mixed with a stoichiometric amount of Mo in accordance with equation XVII, i.e., 60.4% by weight $\text{CH}_6\text{N}_4\text{O}_3$ and 39.6% by weight Mo. The composition autoignited at $230 \pm 2^\circ\text{C}$ by DSC analysis.~~

~~_____ This is an underoxidized reaction which leaves some products in an incompletely oxidized state. If there is an external source of oxygen the reaction proceeds according to equation XVIII:~~



~~_____ This composition points out the utility of using organic nitrates in autoignition reactions.~~

Example 13.



A 1:2 ratio of guanidine nitrate to AgNO_3 was mixed with a stoichiometric amount of Mo in accordance with equation XIX, i.e., 21.9% by weight $\text{CH}_6\text{N}_4\text{O}_3$, 60.9% AgNO_3 and 17.2% by weight Mo. The composition autoignited at $172 \pm 2^\circ\text{C}$ (by DSC).

Q15
Concl'd.

This composition is also an example of organic nitrates in autoignition reactions. However, this composition is fully oxidized, and, therefore, requires no external source of oxygen.
